

Unit 3 Chapter 3 Lessons 1

1. acid ~ a compound that releases hydrogen ions when it is dissolved in water (H^+)
2. strong acid ~ an acid that produces many hydrogen ions when it is dissolved in water and is very corrosive; the three most common strong acids are used in industry (sulfuric acid, nitric acid, and hydrochloric acid) and are good electrolytes
3. weak acid ~ an acid that produces a small number of hydrogen ions when it is dissolved in water and is not very corrosive; examples include acetic acid, carbonic acid, citric acid, and boric acid and are poor electrolytes

Unit 3 Chapter 3 Lessons 2

1. base ~ a compound that releases hydroxide ions when it is dissolved in water (OH^-)
2. strong base ~ a base that produces many hydroxide ions when it is dissolved in water and are good electrolytes; metals in groups 1 and 2 of periodic table form strong bases; examples include sodium hydroxide and calcium hydroxide
3. weak base ~ a base that produces a small number of hydroxide ions when it is dissolved in water and are poor electrolytes; examples include ammonium hydroxide and aluminum hydroxide

Unit 3 Chapter 3 Lessons 3

1. neutralization ~ the reaction between an acid and a base that yields water and often a salt; H^+ ions will equal OH^- ions
2. salt ~ a compound formed from the positive ions of a base and the negative ions of an acid; neutral; examples include sodium chloride (table salt), potassium bromide, ammonium chloride, calcium sulfate, and ammonium nitrate

Unit 3 Chapter 3 Lessons 4

1. indicator ~ a substance that changes color in the presence of an acid or a base; common indicators are litmus dye, phenolphthalein, methyl orange, grape juice, cabbage, and tea
2. pH scale ~ measures the acidity (or alkalinity) of a solution

$\text{pH} > 7$	Basic Solution
$\text{pH} = 7$	Neutral Solution
$\text{pH} < 7$	Acidic Solution